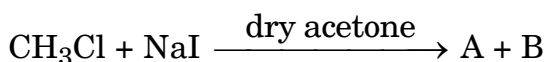


KANKOR EXAMS QUESTIONS

- (A) because they are radioactive in nature
- (B) because they have large atomic numbers
- (C) because they have large atomic masses
- (D) due to comparable energies of 5f, 6d and 7s orbitals

4. Consider the following reaction and identify A and B :



- (A) A = CH₃I, B = NaCl
- (B) A = CH₃OH, B = NaCl
- (C) A = CH₃CHO, B = NaCl
- (D) A = C₂H₆, B = CH₃I

5. The major product of carbylamine reaction is :

- | | |
|---------------------|----------------|
| (A) Carboxylic acid | (B) Aldehyde |
| (C) Cyanide | (D) Isocyanide |

6. Which of the reactions is used in the conversion of a ketone into hydrocarbon ?

- | | |
|-----------------------------|-----------------------------|
| (A) Reimer-Tiemann reaction | (B) Wolff-Kishner reduction |
| (C) Aldol condensation | (D) Stephen reaction |

7. Which of the following ligands forms a chelate complex ?

- (A) Ammonia
- (B) Water
- (C) NO₂⁻
- (D) Oxalate ion





8. Half-life ($t_{1/2}$) of a first order reaction is 1386 s. The value of rate constant is :
- (A) $0.5 \times 10^4 \text{ s}^{-1}$
 (B) $5.0 \times 10^{-4} \text{ s}^{-1}$
 (C) $0.5 \times 10^{-5} \text{ s}^{-1}$
 (D) $0.5 \times 10^{-2} \text{ s}^{-1}$
9. The polypeptide chain in a protein has amino acids linked with each other in a specific sequence. This specific sequence of amino acids is called :
- (A) Primary structure of protein
 (B) Secondary structure of protein
 (C) Tertiary structure of protein
 (D) Quaternary structure of protein
10. Identify the polysaccharide among the following :
- (A) Fructose (B) Maltose
 (C) Glucose (D) Cellulose
11. Consider the following compounds :
 $\text{C}_2\text{H}_5\text{NH}_2$, $(\text{C}_2\text{H}_5)_2\text{NH}$, $\text{C}_6\text{H}_5\text{CH}_2\text{NH}_2$, NH_3 , $\text{C}_6\text{H}_5\text{NH}_2$
 The correct increasing order of the above compounds on the basis of their basic strength is :
- (A) $\text{C}_6\text{H}_5\text{NH}_2 < \text{NH}_3 < \text{C}_6\text{H}_5\text{CH}_2\text{NH}_2 < \text{C}_2\text{H}_5\text{NH}_2 < (\text{C}_2\text{H}_5)_2\text{NH}$
 (B) $\text{NH}_3 < \text{C}_6\text{H}_5\text{NH}_2 < \text{C}_6\text{H}_5\text{CH}_2\text{NH}_2 < \text{C}_2\text{H}_5\text{NH}_2 < (\text{C}_2\text{H}_5)_2\text{NH}$
 (C) $\text{C}_6\text{H}_5\text{CH}_2\text{NH}_2 < (\text{C}_2\text{H}_5)_2\text{NH} < \text{NH}_3 < \text{C}_6\text{H}_5\text{NH}_2 < \text{C}_2\text{H}_5\text{NH}_2$
 (D) $\text{C}_2\text{H}_5\text{NH}_2 < \text{C}_6\text{H}_5\text{NH}_2 < \text{NH}_3 < \text{C}_6\text{H}_5\text{CH}_2\text{NH}_2 < (\text{C}_2\text{H}_5)_2\text{NH}$
12. Primary, secondary and tertiary alcohols can be distinguished by :
- (A) Lucas test
 (B) Fehling's test
 (C) Tollens' test
 (D) Hinsberg's test





For Questions number 13 to 16, two statements are given — one labelled as Assertion (A) and the other labelled as Reason (R). Select the correct answer to these questions from the codes (A), (B), (C) and (D) as given below.

- (A) Both Assertion (A) and Reason (R) are true and Reason (R) is the correct explanation of the Assertion (A).
- (B) Both Assertion (A) and Reason (R) are true, but Reason (R) is **not** the correct explanation of the Assertion (A).
- (C) Assertion (A) is true, but Reason (R) is false.
- (D) Assertion (A) is false, but Reason (R) is true.

13. *Assertion (A) :* All aliphatic aldehydes give a positive Fehling's test.

Reason (R) : Aliphatic aldehydes are reduced by Fehling's reagent.

14. *Assertion (A) :* D (+) – Glucose is dextrorotatory in nature.

Reason (R) : (+) represents dextrorotatory nature and D represents the configuration.

15. *Assertion (A) :* Highest oxidation state of Mn is +7 in first series of transition elements.

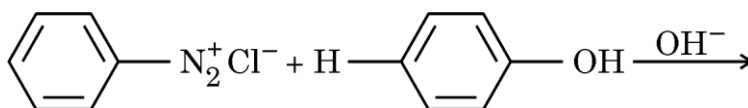
Reason (R) : Transition metals exhibit variable oxidation states.

16. *Assertion (A) :* p-nitrophenol is more acidic than phenol.

Reason (R) : Nitro group is an electron-withdrawing group, it stabilises phenoxide ion by dispersal of negative charge.

SECTION B

17. (a) Complete the following equation :



(b) How will you convert nitromethane to methyl isocyanide ?

1+1





18. (a) Name any two materials other than hydrogen that can be used as fuel in fuel cells.
- (b) Define molar conductivity. 1+1
19. (a) 1.00 molal aqueous solution of trichloroacetic acid is heated to its boiling point. Boiling point of this solution was found to be 100.18°C . Calculate the van't Hoff factor for trichloroacetic acid. (Given : K_b for water = $0.512 \text{ K kg mol}^{-1}$) 2
- OR**
- (b) State Henry's law. Calculate the mole fraction of CO_2 in water at 298 K under 760 mm Hg. (Given : K_H for CO_2 in H_2O at 298 K = $1.25 \times 10^6 \text{ mm Hg}$) 2
20. (a) Write the hydrolysis products of lactose.
- (b) What are non-essential amino acids ? 1+1
21. (a) Name two water-soluble vitamins.
- (b) How will you prove the absence of free – CHO group in the cyclic structure of glucose molecule ? 1+1

SECTION C

22. Calculate emf of the following cell at 298 K : 3
- $\text{Al (s)} | \text{Al}^{3+}(\text{aq}) (0.1 \text{ M}) || \text{Ni}^{2+}(\text{aq}) (0.01 \text{ M}) | \text{Ni (s)}$
- (Given : $E_{\text{Ni}^{2+}/\text{Ni}}^0 = -0.25 \text{ V}$, $E_{\text{Al}^{3+}/\text{Al}}^0 = -1.66 \text{ V}$, $[\log 10 = 1]$)
23. How do you convert the following ? 3
- (a) Toluene to Benzaldehyde
- (b) Benzoyl chloride to Acetophenone
- (c) Bromobenzene to Benzoic acid

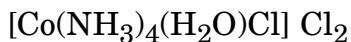




24. The rate of a particular reaction quadruples when the temperature increases from 293 K to 313 K. Calculate the energy of activation (E_a) of the reaction assuming that it does not change with temperature. 3

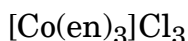
(Given : $\log 4 = 0.60$, $R = 8.314 \text{ JK}^{-1} \text{ mol}^{-1}$)

25. (a) Write the IUPAC name of the given complex : 1



- (b) Define polydentate ligand. Give an example. 1

- (c) Mention the type of isomerism exhibited by the given complex : 1



26. (a) A coordination compound $\text{CrCl}_3 \cdot 6\text{H}_2\text{O}$ is mixed with excess of AgNO_3 solution, one mole of AgCl is precipitated per mole of the compound. Write the structural formula of the coordination compound. 1

- (b) Write the oxidation state and hybridisation of the central metal in the given complex : $[\text{Fe}(\text{CN})_6]^{3-}$. [Atomic number : $\text{Fe} = 26$] 1

- (c) Why is $[\text{Ti}(\text{H}_2\text{O})_6]^{3+}$ coloured ? [Atomic number : $\text{Ti} = 22$] 1

27. (a) Define order of a reaction.

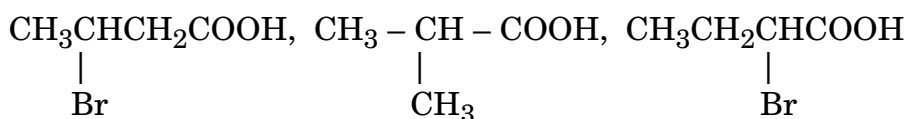
- (b) The rate for the following reaction is given by : $\text{A} + \text{B} \rightarrow \text{C}$

$$\text{Rate} = k[\text{A}][\text{B}]^2$$

- (i) How is the rate of reaction affected if we double the concentration of B ?

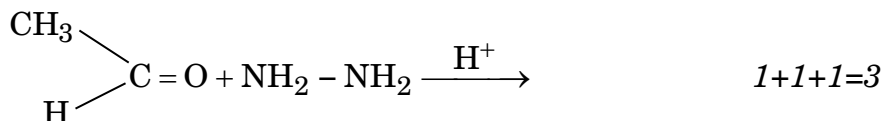
- (ii) Write the overall order of a reaction if 'A' is present in large excess. 1+2=3

28. (a) (i) Arrange the following compounds in increasing order of their acidic strengths :





- (ii) Why is CH_3CHO more reactive than acetone towards reaction with HCN ?
- (iii) Complete the equation :



OR

- (b) An organic compound with the molecular formula $\text{C}_8\text{H}_8\text{O}$ forms 2,4-DNP derivative, reduces Tollens' reagent and undergoes Cannizzaro reaction. On vigorous oxidation it gives Benzene-1,2-dicarboxylic acid. Identify the compound and write the reactions of compound 2,4-DNP and when it undergoes Cannizzaro reaction.

3

SECTION D

The following questions are Case-based questions. Read the case carefully and answer the questions that follow.

- 29.** Electrochemistry is the study of the relationship between chemical energy and electrical energy. Many spontaneously occurring chemical reactions liberate electrical energy. In electrolysis, electrical energy is converted directly into chemical energy. The product of an electrolytic reaction depends on the nature of the material being electrolysed and the type of electrode used. Oxidising and reducing species present in the electrolytic cell and their standard electrode potential too, affect the products of electrolysis. Electrolysis plays an important role in most people's daily lives, whether it is for the manufacturing of aluminium, electroplating of metals, or the synthesis of chemical compounds.

Michael Faraday was the first scientist who proposed two laws to explain the quantitative aspects of electrolysis, popularly known as Faraday's laws of electrolysis. Faraday's laws of electrolysis provide a basis for mathematical analysis of the mass deposited at electrodes and the amount of charge passed through them.

Faraday's laws are fundamental in various applications, including electroplating, metal extraction, battery technology and chemical synthesis. These laws also help in environmental monitoring and in various chemistry experiments.



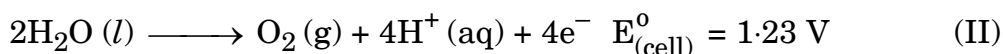
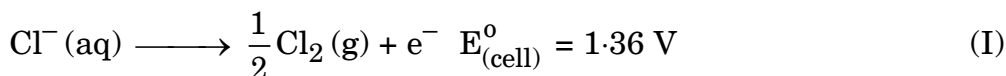


Answer the following questions :

- (a) Predict the products of electrolysis in each of the following : 2
- (i) An aqueous solution of CuCl_2 with platinum electrodes.
- (ii) A concentrated solution of H_2SO_4 with platinum electrodes.
- (b) (i) How much charge in faraday is required for the reduction of 1 mol of Ag^+ to Ag ? 1

OR

- (b) (ii) State Faraday's second law of electrolysis. 1
- (c) The following reactions occur at the anode during the electrolysis of aqueous sodium chloride solution :



Which reaction is feasible at the anode and why ? 1

30. Ethers are prepared by the dehydration of alcohols in presence of protic acids at 413 K. Symmetrical and unsymmetrical ethers can also be prepared by Williamson synthesis. This reaction involves $\text{S}_{\text{N}}2$ attack of an alkoxide ion on primary alkyl halide. If a tertiary alkyl halide is used, elimination reaction occurs and an alkene is formed, no ether is formed. C – O bond in ethers are cleaved under drastic conditions with excess of HI. When unsymmetrical ethers react with HI, the alkyl halide is formed from smaller alkyl group. If one of the alkyl group is tertiary, the alkyl halide is formed from the tertiary alkyl group because tertiary alkyl carbocation is more stable than the primary carbocation. Cleavage of alkyl aryl ethers takes place at the alkyl-oxygen bond due to more stable aryl-oxygen bond. The order of reactivity of hydrogen halides is $\text{HI} > \text{HBr} > \text{HCl}$. Aromatic ethers undergo electrophilic substitution reactions. The alkoxy group attached to the aromatic ring activates the ring towards electrophilic substitution and directs the incoming group to ortho- and para-positions.

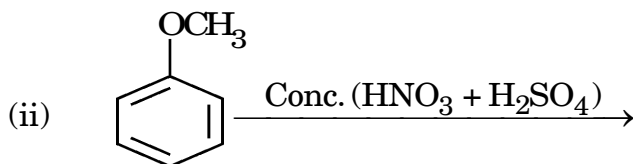
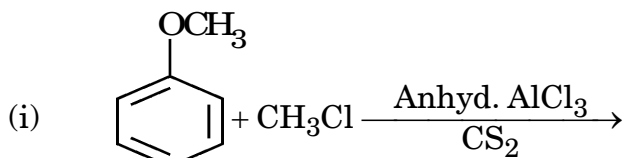




Answer the following questions :

(a) Complete the following equations :

2



(b) (i) Write the names of alkyl halide and sodium alkoxide used to prepare tert-butyl ethyl ether.

1

OR

(b) (ii) Anisole on reaction with HI gives phenol and $\text{CH}_3 - \text{I}$ and not methanol and iodobenzene. Justify the statement.

1

(c) Why is $\text{C} - \text{O} - \text{C}$ bond angle in ethers slightly greater than tetrahedral angle ?

1

SECTION E

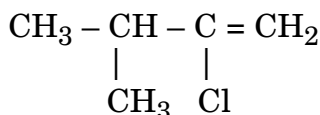
31. (a) (i) (I) Which of the following is more reactive towards $\text{S}_{\text{N}}1$ reaction :

1

2-Bromo-2-methylbutane or 1-Bromopentane

(II) What type of halide is present in the following compound :

1



(III) Why is chloroform stored in dark coloured bottles ?

1

(ii) Define the following terms :

1+1

(I) Ambident Nucleophiles

(II) Racemic mixture

OR

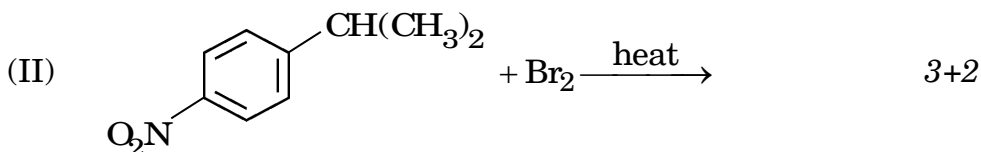
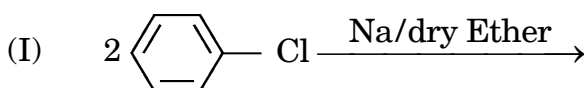




(b) (i) Answer the following :

- (I) Which isomer of C_4H_9Br is most reactive towards S_N1 reaction ?
- (II) Predict the alkene that would be formed by dehydrohalogenation of 1-Bromo-1-methylcyclohexane.
- (III) Although chlorine shows strong $-I$ effect, yet it is ortho/para-directing in electrophilic aromatic substitution reactions. Why ?

(ii) Write the major product in the following reactions :



32. (a) (i) Calculate the freezing point of a solution when 10.5 g of MgBr_2 was dissolved in 250 g of water, assuming MgBr_2 undergoes complete dissociation. 3
- (Given : Molar mass of $\text{MgBr}_2 = 184 \text{ g mol}^{-1}$,
 K_f for water = $1.86 \text{ K kg mol}^{-1}$)

(ii) Write two differences between ideal and non-ideal solutions. 1+1

OR

- (b) (i) A solution is prepared by dissolving 0.025 g of potassium sulphate in 2 L of water at 27°C . Assuming potassium sulphate is completely dissociated, determine its osmotic pressure. 3
- (Given : $R = 0.082 \text{ L atm K}^{-1} \text{ mol}^{-1}$,
 Molar mass of $\text{K}_2\text{SO}_4 = 174 \text{ g mol}^{-1}$)
- (ii) What type of azeotrope will be formed by a solution of acetone and chloroform ? Give reason. 2





- 33.** (a) (i) (I) Why do transition metals show variable oxidation states ? 1
- (II) Out of Mn^{2+} and Ti^{2+} which will be more paramagnetic and why ? [Atomic No. : Ti = 22, Mn = 25] 1
- (III) Which ion is the strongest oxidising agent in the options given below :
 Cr^{3+} , V^{3+} , Mn^{3+}
 Give reason. 1
 [Atomic No. : Cr = 24, V = 23, Mn = 25]
- (ii) Complete and balance the following equations : 1+1
- (I) $2\text{MnO}_2 + 4\text{KOH} + \text{O}_2 \longrightarrow$
- (II) $5\text{C}_2\text{O}_4^{2-} + 2\text{MnO}_4^- + 16\text{H}^+ \longrightarrow$
- OR**
- (b) (i) What is meant by lanthanoid contraction ?
- (ii) Why do transition metals form coloured compounds ?
- (iii) Why are $E_{\text{M}^{2+}/\text{M}}^0$ values for Mn and Zn more negative than expected ?
- (iv) Which is the most stable oxidation state of Cu and why ?
- (v) Why is Ce^{4+} in aqueous solution a good oxidising agent ? 5×1=5

